Int. J. Heat Mass Transfer. Vol. 12, pp. 989–990. Pergamon Press 1969. Printed in Great Britain

ERRATA

L. P. FILIPPOV: Research of liquid thermal conductivity at Moscow University, *Int. J. Heat Mass Transfer* **11**, 331–345 (1968).

The printer regrets that in Table 2 (p. 337) some figures have been misplaced. The table should begin as follows:

Table 2

Formula	Substances	Temperature range		kcal/mh degC	$\alpha 10^3$
		lower	upper		deg C^{-1}
CCl_4	carbon tetrachloride	15	90	$0{\cdot}087_5$	2
$CHBr_3$	bromoform	10	90	0·087	2·1
$CHCl_3$	chloroform	15	70	0·101	1·8
CH_2O_2	formic acid	15	90	$0{\cdot}23_5$	—
CH_4O	methanol	15	90	0·175	1
$C_2H_4Br_2$	ethylen bromide (1·2-ethylene dibromide)	15	90	$0{\cdot}086_5$	1·3
$C_2H_4Cl_2$	ethylen dichloride (1·2-ethylene dichloride)	15	80	$0{\cdot}115_5$	1·5
C_2H_4O	acetaldehyde	10	30	$0{\cdot}159_5$	$2{\cdot}_6$
$C_2H_3ClO_2$	monochloracetic acid (supercooled)	40	90	0·145	non-linear
$C_2H_4O_2$	acetic acid	5	90	0·143	1
$C_2H_4O_2$	methyl-formate	10	40	0·163	1·2
C_2H_5Br	ethyl-bromide	15	45	$0{\cdot}087_5$	2
C_2H_5I	ethyl-iodide	20	70	$0{\cdot}077_5$	$1{\cdot}_0$
C_2H_5ON	acetamide	65	100	0·216	~0
C_2H_6O	ethanol	15	90	0·140	$1{\cdot}_9$
$C_2H_6O_2$	ethylen-glycol	15	90	0·220	−1·0
C_3H_6O	acetone	15	50	$0{\cdot}139_5$	$2{\cdot}_2$
$C_3H_6O_2$	ethyl-formate	15	50	0·139	(3)
$C_3H_6O_2$	methyl-acetate	10	50	0·134	2
C_3H_8O	propanol	15	80	0·13	1·4
C_3H_8O	*i*-propanol	10	80	0·120	1
$C_3H_8O_3$	glycerine	15	90	0·244	$-2{\cdot}_3$
$C_4H_6O_2$	acetaldehyde	15	90	$0{\cdot}140_5$	1·5
$C_4H_8O_2$	butyric acid	10	90	0·128	0·9
$C_4H_8O_2$	ethyl-acetate	15	90	$0{\cdot}123_5$	2
$C_4H_8O_2$	propyl-formate	15	70	$0{\cdot}124_5$	1·7
$C_4H_{10}O$	butanol	15	90	0·123	1·4
$C_4H_{10}O$	*i*-butanol	15	90	0·115	1·0
$C_4H_{10}O$	diethylether	10	70	0·114	$2{\cdot}_3$

J. W. ROSE: Condensation of a vapour in the presence of a non-condensing gas, *Int. J. Heat Mass Transfer* **12**, 233–237 (1969).
In equation (8) on p. 234 (W^0 should read W_0).